

Naming Ionic Compounds 1

Part 1

Write the formula for each of the following ionic compounds.

1. lithium acetate Li CH₃COO
2. iron(II) phosphate Fe₃(PO₄)₂
3. iron(III) phosphate Fe PO₄
4. titanium(II) selenide Ti Se
5. calcium bromide Ca Br₂
6. gallium(III) chloride Ga Cl₃
7. sodium hydride Na H
8. beryllium hydroxide Be(OH)₂
9. zinc carbonate Zn CO₃
10. manganese(VII) arsenide Mn₃As₇
11. copper(II) chlorate Cu(ClO₃)₂
12. cobalt(III) chromate Co₂(CrO₄)₃
13. ammonium oxide (NH₄)₂O
14. potassium hydroxide KOH
15. lead(IV) sulfate Pb(SO₄)₂
16. silver cyanide Ag CN
17. vanadium(V) nitride V₅N₅
18. strontium acetate Sr(CH₃COO)₂
19. platinum(II) sulfide PtS
20. ammonium sulfate (NH₄)₂SO₄

40

20

Part 2

Name each of the following ionic compounds. (Hint: 33 – 40 require roman numerals)

21. NaBr sodium bromide
22. Sc(OH)_3 scandium hydroxide
23. NH_4F ammonium fluoride
24. CaCO_3 calcium carbonate
25. Na_2SO_4 sodium sulfate
26. Li_2SO_3 lithium sulfite
27. Zn_3P_2 zinc phosphide
28. $\text{Sr(C}_2\text{H}_3\text{O}_2)_2$ strontium acetate
29. Ag_3PO_4 silver phosphate
30. YClO_3 yttrium chlorate
31. CdSO_3 cadmium sulfite
32. KMnO_4 potassium permanganate
33. Pb_3N_2 lead(II) nitride
34. CoCO_3 cobalt(II) carbonate
35. $\text{Cu(NO}_2)_2$ copper(II) nitrite
36. $\text{Fe(HCO}_3)_2$ iron(II) bicarbonate
37. $\text{V}_2(\text{SO}_4)_3$ vanadium(III) sulfate
38. Cu_2O copper(I) oxide
39. SnS_2 tin(IV) sulfide
40. Ti(CN)_4 titanium(IV) cyanide